Curiosity: A Fizzy Drink Is Not a Frozen System

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An unopened carbonated drink can look completely still, yet carbon dioxide molecules are continuously moving between the liquid and the small gas space above it. A dynamic-equilibrium model asks about two directions: molecules entering the liquid and molecules leaving it. When those average rates match under stated conditions, the visible amounts can remain steady even though molecular events continue.

Opening the container changes the conditions, especially the gas pressure above the liquid. Bubbles then grow and escape. The model is useful because it distinguishes a steady observation from a stopped process; it is not a claim that every beverage, temperature, or container has the same equilibrium constant.

For this packet, keep the context conceptual. Do not seal containers or attempt to create gas pressure. The supplied event cards and concentration models are sufficient evidence for the equilibrium reasoning target.